

N1

 Ne, F_2, Ar, Cl_2

$$m(Ne) = 1 \mu / 22,4 \mu / \text{моль} \cdot 20 \mu / \text{моль} = 0,892$$

$$m(F_2) = 1 \mu / 22,4 \mu / \text{моль} \cdot 38 \mu / \text{моль} = 1,692$$

$$m(Cl_2) = 1 \mu / 22,4 \mu / \text{моль} \cdot 71 \mu / \text{моль} = 3,172$$

$$m(Ar) = 1 \mu / 22,4 \mu / \text{моль} \cdot 40 \mu / \text{моль} = 1,792$$

N2

$$1. \text{ алюмин } V(Au) = 12 / 19,30 \mu / \text{см}^3 = 0,05 \text{ см}^3$$

$$2. \text{ мосс } V(Cu) = 12 / 8,92 \mu / \text{см}^3 = 0,11 \text{ см}^3$$

$$3. \text{ темир } V(Fe) = 12 / 7,872 \mu / \text{см}^3 = 0,13 \text{ см}^3$$

$$4. \text{ титан } V(Ti) = 12 / 4,52 \mu / \text{см}^3 = 0,22 \text{ см}^3$$

N3

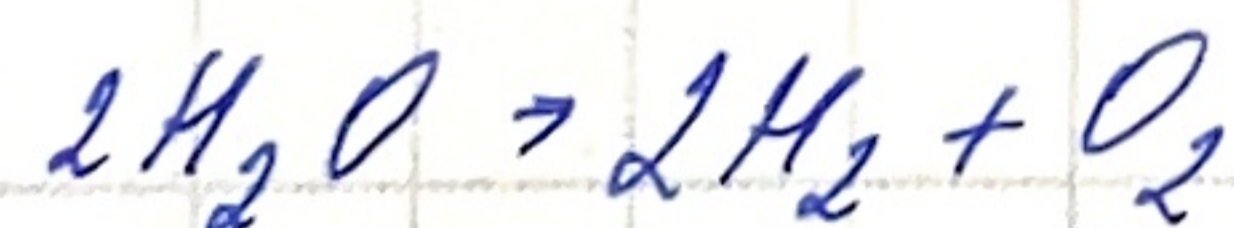
$$1. CaO \quad M_r(CaO) = 40 + 16 = 56$$

$$2. HNO_3 \quad M_r(HNO_3) = 1 + 14 + 16 \cdot 3 = 63$$

$$3. Ca(OH)_2 \quad M_r(Ca(OH)_2) = 40 + 16 + (1 \cdot 2) = 74$$

$$4. CaSO_4 \quad M_r(CaSO_4) = 40 + 32 + (16 \cdot 4) = 136$$

N4



$$n = V / 22,4 = 25000 / 22,4 = 1116 \text{ моль}$$

$$m = M \cdot n = 65 \cdot 1116 = 72540 \text{ г}$$